Answer all of the following questions. Remember to show all work and pay attention to significant figures and units.

1. Why is chewing gum forbidden in the chemistry laboratory?
   
   The gum could become contaminated by chemicals and poison you.

2. A particular experiment calls for about 1.5 g of sodium chloride. You accidentally weigh out over 3 g of NaCl. What should you do in this situation?
   
   a. Dump all of the NaCl back into the original container and weigh it out again using less.
   b. Put a little of the NaCl that you weighed out back into the original container so that you end up with the correct amount on the balance.
   c. Use all of the NaCl that you weighed out even though it is a lot more than the experiment calls for.
   d. Remove a little of the NaCl that you weighed out (so that you end up with the correct amount on the balance) and dispose of it in the trash.

3. Which one of the following volume measuring devices gives a volume reading of 34.75 mL?

4. The measuring device with the correct reading in #3 is a (circle one) buret graduated cylinder.

5. Percent relative error can be calculated by using the formula

   \[
   \% \text{ error} = \frac{(\text{experiment} - \text{accepted value})}{\text{accepted value}} \times 100\%.
   \]

   If a student determines that the percent tin in an unknown sample is 34.8%, but the actual accepted value for the percent tin is 35.89%, what is the percent error in the student's result? Remember to show all of your work and pay attention to units and significant figures.

   \[
   \% \text{ error} = \frac{(34.8\% - 35.89\%)}{35.89\%} \times 100\% = -3.07\%.
   \]

   Version 2: \[
   \frac{(33.8\% - 34.87\%)}{34.87\%} \times 100\% = -3.17\%.
   \]